



HAL
open science

Thermodynamics of aqueous solutions at high ionic strength using the Pitzer model.

Arnault Lassin, Adeline Lach, Laurent André

► **To cite this version:**

Arnault Lassin, Adeline Lach, Laurent André. Thermodynamics of aqueous solutions at high ionic strength using the Pitzer model. . International Symposium on Solubility Phenomena and Related Equilibrium Processes (ISSP) - 2018, Jul 2018, Tours, France. hal-01720540

HAL Id: hal-01720540

<https://brgm.hal.science/hal-01720540>

Submitted on 1 Mar 2018

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Abstract submitted for: Oral and Paper ; Oral ; Flash Oral; Poster

THERMODYNAMICS OF AQUEOUS SOLUTIONS AT HIGH IONIC STRENGTH USING THE PITZER MODEL

A. Lassin¹, A. Lach¹, L. André^{1,2}

¹BRGM, Orléans, France

²ISTO UMR 7327 Université d'Orléans-CNRS-BRGM, Orléans, France

e-mail (corresponding author): a.lassin@brgm.fr

More and more attention is paid to the description of brine chemistry for various reasons. Natural brines are extensively exploited for the reserves of valuable chemical elements or chemical products they contain (such as lithium, potash, potassium, magnesium and other various salts). Geothermal energy currently uses hot deep brines. Greenhouse gases like CO₂ are forecasted to be stored in deep geological reservoirs containing saline solutions. The number of desalination plants is increasing with the fresh water demand, leading to increasing amounts of highly saline waste.

Geochemical modelling is one of the numerical tool needed to predict the properties and the behavior of such systems. Different approaches can be used but one of the most relevant is the Pitzer¹ model.

Since more than 20 years, our team is working on developing and improving the modelling of multicomponent saline systems, using the Pitzer approach. Our efforts focus on both the development of a database for specific interaction parameters and a numerical code. We propose a Pitzer database for several chemical systems in order to deal with lithium chemistry^{2,3}, phosphate systems, chloride systems (Na⁺, K⁺, Ca⁺² and Mg⁺²)-Cl between 0-100°C⁴, and nitrate-sulfate systems⁵ at 25°C, both in acidic and basic conditions. These implementations in the database allow dealing with topics such as the optimisation of lithium extraction process from salars brines or the improvement of phosphate production from WPPA process. Next evolution perspectives of this database are the extension to high temperatures (above 150°C) and the addition of trace elements (aluminium and radionuclides for instance).

The description of these complex aqueous solutions requires the development of specific tools, like PhreeSCALE⁶, after the geochemical code PHREEQC⁷. In this new tool, computation of heat capacities, excess enthalpy and density was implemented in coherence with Pitzer and HKF⁸ models. Consequently, the heat capacity or density of a geothermal fluid can be computed according to its exact chemical composition.

References:

- (1) Pitzer, K. S. **1991** *Activity Coefficients in Electrolyte Solutions*
- (2) Lassin, A.; et al. *Am. J. Sci.* **2015**, *315*, 204–256 10.2475/03.2015.02.
- (3) Lassin, A.; et al. *CALPHAD* accepted.
- (4) Lach, A.; et al. *J. Chem. Eng. Data* **2017**, *62* (10), 3561–3576 10.1021/acs.jced.7b00553.
- (5) Lach, A.; et al. *J. Chem. Eng. Data* **2018**, in press 10.1021/acs.jced.7b00953.
- (6) Lach, A.; et al. *Comput. Geosci.* **2016**, *92*, 58–69 10.1016/j.cageo.2016.03.016.
- (7) Parkhurst, D. L.; Appelo, C. A. J. **2013** *Description of Input and Examples for PHREEQC Version 3*
- (8) Helgeson, H. C.; et al. *Am. J. Sci.* **1981**, *281*, 1249–1516 10.2475/ajs.281.10.1249.

